

# Chapter 9 2 Stoichiometry

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## Chapter 9 2 Stoichiometry

9-2 Ideal Stoichiometric Calculations Ideal Stoichiometry - All reactants are converted into products I. A Common Method for Solving All Stoichiometry Problems A. Mass-Mass Problems 1. Start with a known mass of reactant or product, find an unknown mass of another reactant or product 2. All other stoichiometry problems are derivations ...

## Chapter 9 - Stoichiometry

Stoichiometry is the part of chemistry that applies the balanced chemical equation to determine the quantities of reactants and products. Interpreting balanced equations. ... Chapter 9: Section 2: Ideal Stoichiometric Calculations Last modified by: Michelle Stover ...

## Chapter 9: Section 2: Ideal Stoichiometric Calculations

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## Chapter 9 Stoichiometry | Academic

290 Chapter 9 DO NOT EDIT--Changes must be made through "File info" CorrectionKey=NL-A. Critical Thinking Reviewing Main Ideas 1. What is stoichiometry? 2. For each equation, write all possible mole ratios. a.  $2\text{HgO}(s) \rightarrow 2\text{Hg}(l) + \text{O}_2(g)$  b.  $4\text{NH}_3(g) + 6\text{NO}(g) \rightarrow 5\text{N}_2(g) + 6\text{H}_2\text{O}(g)$

## CorrectionKey=NL-A DO NOT EDIT--Changes must be made ...

The four quantities involved in stoichiometric calculations are: • particles - the relative amounts of atoms, ions, unit formulas or molecules in various reactants or products • moles - the relative number of moles of reactants or products • mass - the relative masses of reactants or products • volume - the relative amounts of gaseous reactants or products Atoms and mass are always conserved in chemical reactions.

## CHEMISTRY NOTES - Chapter 9 Stoichiometry

CHAPTER 9 REVIEW. Stoichiometry. SECTION 9.2. PROBLEMS Write the answer on the line to the left. Show all your work in the space provided. 1. The following equation represents a laboratory preparation for oxygen gas:  $2\text{KClO}_3(s) \rightarrow 2\text{KCl}(s) + 3\text{O}_2(g)$  How many grams of  $\text{O}_2$  form if 3.0 mol of  $\text{KClO}_3$  are totally consumed? 2. Given the following equation ...

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## Quizlet

CHAPTER 9 REVIEW Stoichiometry SECTION 3 PROBLEMS Write the answer on the line to the left. Show all your work in the space provided. 1. 88% The actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage yield. 2. 6.0 mol of N<sub>2</sub> are mixed with 12.0 mol of H

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CHAPTER 9 REVIEW. Stoichiometry. MIXED REVIEW. SHORT ANSWER Answer the following questions in the space provided. 1. Given the following equation:  $C_3H_4(g) + x O_2(g) \rightarrow 3CO_2(g) + 2H_2O(g)$  a. What is the value of the coefficient  $x$  in this equation? b. What is the molar mass of  $C_3H_4$ ? c. How many moles are in an 8.0 g sample of  $C_3H_4$ ? 2. a. What ...

## CHAPTER 9 REVIEW

Chapter 9 9.2 Objectives • Calculate the amount in moles of a reactant or a product from the amount in moles of a different reactant or product. • Calculate the mass of a reactant or a product from the amount in moles of a different reactant or product.

## Chapter 9 Stoichiometry

stoichiometry (which you studied in Chapter 3) deals with the mass relationships of elements in compounds. Reaction stoichiometry involves the mass relationships between reactants and products in a chemical reaction. Reaction stoichiometry is the subject of this chapter and it is based on

## CHAPTER 9 Stoichiometry

Stoichiometry Chapter 9, p. 275 - 294 Intro to Stoichiometry • Reaction Stoichiometry: Involves the mass relationships between reactants and products in a chemical reaction • Coefficients in a chemical reaction represent the mole ratios of each substance that react together • Example:  $4Fe + 3O_2 \rightarrow 2Fe_2O_3$

## Chapter 9 Stoichiometry Notes - Chemistry

Define stoichiometry: 299 Write a mole ratio for a balanced chemical equation and tell to importance of a mole ratio.: 299—301 Use molar ratios and molar masses to create

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conversion factors for solving stoichiometric problems.:299—311 Distinguish between a limiting reactant (reagent) and an excess reactant (reagent).:312—315 Identify the limiting reactant in a problem and calculate the ...

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### **Chapter 9 - Stoichiometry - Ms. Clark's Website**

Steps for Stoichiometry: 1- Identify the given and target compound 2-Balance the equation for the reaction 3- Set up the problem (convert to moles if necessary) 4-Use the mole ratio(s) to calculate the number of moles of the desired compound 5- convert to grams of the desired compound if necessary

### **Chapter 9: Stoichiometry - J.G.M.C.K.**

5a. 2 N<sub>2</sub>O plus 3 O<sub>2</sub> yields 4 NO<sub>2</sub>. b. 4 NO<sub>2</sub> to 3 O<sub>2</sub>. c. The ratio between NO<sub>2</sub> and O<sub>2</sub> in the balanced equation is 4 to 3. so you will have the same ratio of 20 moles NO<sub>2</sub> to 15 moles O<sub>2</sub>. d. true. e. Now look at the ratio between total atomic wts for 2N<sub>2</sub>O and 4 NO<sub>2</sub>. 2N<sub>2</sub>O adds up to 88. 4 NO<sub>2</sub> adds up to 128 so this is false.

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