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Kinetic Molecular Theory ©POGIL – 2005 5/5 Authored by Applications 1. There is a government warning on all aerosol cans that states: Do not store at a temperature above 120° F (50°C). a) Explain why this warning is required in terms of the relationship between temperature and pressure and the kinetic molecular theory.

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Kinetic Molecular Theory Pogil 2005 Answers
Kinetic Molecular Theory ©POGIL – 2005 4/4 Authored by Key Questions 1. What causes a gas to exert pressure when confined in a Page 8/23. Bookmark File PDF Pogil Kinetic Theory Answer Key container? 2. How does the total volume of gas particles compare to the volume of the space between the gas particles? 3. As the

Pogil Kinetic Theory Answer Key - Aplikasi Dapodik
POGIL: Kinetic Molecular Theory. Modified from Foundations of Chemistryby David HansonPage 1 of 3. POGIL: Kinetic Molecular Theory. Learning Objectives. Identify the basic differences between particle behavior in the solid, liquid, and gaseous phases. Develop an understanding of the postulates of the kinetic molecular theory. Identify how temperature affects molecular motion. Apply the kinetic molecular theory to predict the outcome of everyday situations.

POGIL: Kinetic Molecular Theory
The kinetic molecular theory (KMT) is a simple microscopic model that effectively explains the gas laws described in previous modules of this chapter. This theory is based on the following five postulates described here. (Note: The term "molecule" will be used to refer to the individual chemical species that compose the gas, although some ...

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The average kinetic energy of the gas particles is directly proportional to the Kelvin temperature of the gas. 3/3 - 2005 Authored by Edited by Linda Padwa and David Hanson, Stony Brook University Kinetic Molecular Theory Key Questions 1. What causes a gas to exert pressure when confined in a container? 2.

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Kinetic Molecular Theory Stony Brook Answers [EBOOK]
The particles are assumed to not attract nor repel each other. The average kinetic energy of the gas particles is directly proportional to the Kelvin temperature of the gas. 3/3 ©POGIL -2005 Authored by Edited by Linda Padwa and David Hanson, Stony Brook University. Kinetic Molecular Theory Key Questions 1.

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9.5 The Kinetic-Molecular Theory – Chemistry
A. Kinetic Molecular Theory (KMT) = the idea that particles of matter are always in motion and that this motion has consequences. 1) theory developed in the late 19th century to account for the behavior of the atoms and molecules that make up matter 2) based on the idea that particles in all forms of matter are

I. MOLECULES IN MOTION: A.
The average kinetic energy of the gas particles is directly proportional to the Kelvin temperature of the gas. 3/3 ©POGIL - 2005 Authored by Edited by Linda Padwa and David Hanson, Stony Brook University

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• Collisions between molecules or the container walls are elastic; i.e., no loss of kinetic energy or momentum. • Gas pressure arises from molecules striking the walls of the container. • The average kinetic energy is proportional to the absolute temperature.

Chem 116 POGIL Worksheet - Week 2 Gas Laws - Part 2
POGIL: Kinetic Molecular Theory - Studyres The Kinetic-Molecular Theory Explains the Behavior of Gases, Part II According to Graham's law, the molecules of a gas are in rapid motion and the molecules themselves are small. The average distance between the molecules of a gas is large compared to the size of the molecules. Page 1/5

Kinetic Molecular Theory Pogil Answer
Kinetic Molecular Theory states that gas particles are in constant motion and exhibit perfectly elastic collisions. Kinetic Molecular Theory can be used to explain both Charles' and Boyle's Laws. The average kinetic energy of a collection of gas particles is directly proportional to absolute temperature only.

Kinetic Molecular Theory | Boundless Chemistry
The Kinetic Molecular Theory and Graham's Laws. The kinetic molecular theory can be used to explain the results Graham obtained when he studied the diffusion and effusion of gases. The key to this explanation is the last postulate of the kinetic theory, which assumes that the temperature of a system is proportional to the average kinetic energy ...

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